

GOVERNMENT DEGREE COLLEGE
RAMPACHODAVARAM
ALLURI SEETHA RAMA RAJU DISTRICT
DEPARTMENT OF CHEMISTRY

QUIZ

2017-2022

<i>YEAR</i>	<i>NO.OF QUIZ</i>	<i>DATE</i>	<i>CLASS</i>	<i>TOPIC</i>
2017-18	1	20-02-2018	I,II&IIIMPC	Chemical bonding
2018-19	1	15-02-2019	I,II,&IIIMPC ,BZC	Group-13 Elements
2019-20	1	26-02-2020	I,II,&III MPC ,BZC	Carbohydrates
2020-21	1	25-02-2021	I,II,&III MPC,BZC	Group-14 Elements
2021-22	1	17-02-2022	I,II,&III MPC,BZC	General Chemistry

GOVERNMENT DEGREE COLLEGE
RAMPACHODAVARAM
ALLURI SEETHA RAMA RAJU DISTRICT

DEPARTMENT OF CHEMISTRY

QUIZ COMPETITION FOR I ,II,III B.Sc

2017-2018

GOVERNMENT DEGREE COLLEGE
RAMPACHODAVARAM
DEPARTMENT OF CHEMISTRY

REPORT ON STUDENT QUIZ –I,II,III- B.Sc

The Department of Chemistry Conducted a student Seminar for the students of II B.Sc mpc Students on 20-02-2018. The topic For the QUIZ was Chemical Bonding .

Majority of the students from I B.Sc II B.Sc III B.Sc were attended for the QUIZ .The topic was enlightened and informative to the students.

Few Students have asked questions related to the topic .After the completion of seminar feed back was also Collected from the students.

TO
THE PRINCIPAL,
GOVT.DEGREE COLLEGE,
RAMPACHODAVARAM.

RAMPACHODAVARAM,
10-02-2018.

Respected Sir,

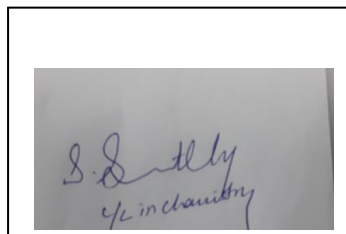
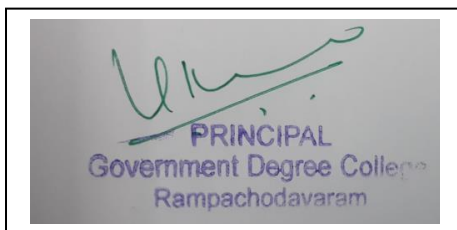
SUB: Seeking permission to conduct student seminars for the year 2017-18 Req reg.

I S.Swarna Latha,Lecturer in chemistry,Govt.Degree College, Rampachodavaram, planned to conduct student seminars for the year 2017-18 for I,II,III B.Sc[MPC] classes.

So kindly grant me permission to conduct seminars.

Thanking you sir.

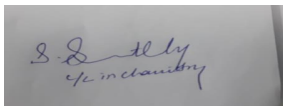
Yours faithfully



GOVERNMENT DEGREE COLLEGE
RAMPACHODAVARAM
DEPARTMENT OF CHEMISTRY

CIRCULAR

The Department of Chemistry is going to be conducted a student Quiz on 20-02-2018. All the students must attend the Quiz in The chapter Chemical bonding Chemistry Lab Room:17 without fail

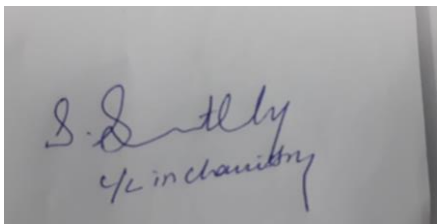
A rectangular box containing a handwritten signature in black ink. The signature appears to be 'S. Swarna Latha' with 'c/o. in chemistry' written below it.

S.Swarna Latha

Lecturer in Chemistry

GOVERNMENT DEGREE COLLEGE
RAMPACHODAVRAM
DEPARTMENT OF CHEMISTRY
REPORT ON QUIZ PROGRAMME

The Department Of Chemistry has conducted an inter group quiz programme on 20-02-2018. It was a 3 stage competition which covered the areas of In-Organic, Organic and Physical chemistry. There were a total 4 teams. Every team had six members from first, second and third year of students. The programme was started at 11:00AM. The Judging panel consists of two judges Sri JEJIBABU Lecturer in COMMERCE and G. LAKSHMI KUMARI LECTURER in Political Science. In the first round the questions were asked from the in-Organic chemistry The winners of this round was team A. Second round questions were asked from Organic chemistry. The winners of this round was team C. In the third round the competition. Over all this inter group quiz competition was a huge triumph. College principal Dr.V.Srinivasa Rao congratulated the winners.



S. Senthil Kumar
H. in Chemistry

Quiz

1. Isostructural species are those which have the same shape and hybridisation. Among the given species identify the isostructural pairs.

(i) [NF₃ and BF₃]

(ii) [BF₄⁻ and NH₄⁺]

(iii) [BCl₃ and BrCl₃]

(iv) [NH₃ and NO₃⁻]

2. Polarity in a molecule and hence the dipole moment depends primarily on electronegativity of the constituent atoms and shape of a molecule. Which of the following has the highest dipole moment?

(i) CO₂

(ii) HI

(iii) H₂O

(iv) SO₂

3. The types of hybrid orbitals of nitrogen in NO⁺, NO⁻ and NH₂⁺ respectively are

expected to be

(v) *sp*, *sp*³ and *sp*²

(vi) *sp*, *sp*² and *sp*³

(vii) *sp*², *sp* and *sp*³

(viii) *sp*², *sp*³ and *sp*

4. Hydrogen bonds are formed in many compounds e.g., H₂O, HF, NH₃. The boiling point of such compounds depends to a large extent on the strength of hydrogen bond and the number of hydrogen bonds. The correct decreasing order of the boiling points of above compounds is :

(ix) HF > H₂O > NH₃

(x) H₂O > HF > NH₃

(xi) NH₃ > HF > H₂O

(xii) NH₃ > H₂O > HF

5. In PO₄³⁻ ion the formal charge on the oxygen atom of P–O bond is

(i) + 1

(ii) 1

(iii) - 0.75

(iv) + 0.75

6. In NO⁻ ion, the number of bond pairs and lone pairs of electrons on nitrogen atom are

(i) 2, 2

(ii) 3, 1

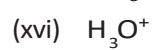
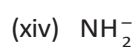
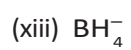
(iii) 1, 3

(iv) 4, 0





7. Which of the following species has tetrahedral geometry?



8. Number of π bonds and σ bonds in the following structure is—

(i) 6, 19

(ii) 4, 20

(iii) 5, 19

(iv) 5, 20

9. Which molecule/ion out of the following does **not** contain unpaired electrons?

(xvii) N_2^+

(xviii) O_2

(xix) O_2^{2-}

(xx) B_2

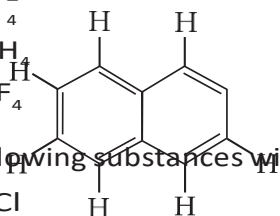
10. In which of the following molecule/ion all the bonds are **not** equal?

(xxi) XeF_4

(xxii) BF_4^-

(xxiii) C_2H_2

(xxiv) SiF_4



11. In which of the following substances will hydrogen bond be strongest?

(xxv) HCl

(xxvi) H_2O

(xxvii) HI

(xxviii) H_2S

12. If the electronic configuration of an element is $1s^2 2s^2 2p^6 3s^2 3p^6 3d^2 4s^2$, the four electrons involved in chemical bond formation will be _____.

(xxix) $3p^6$

(xxx) $3p^6, 4s^2$

(xxxi) $3p^6, 3d^2$

(xxxii) $3d^2, 4s^2$

13. Which of the following angle corresponds to sp^2 hybridisation?

(i) 90°

(ii) 120°

(iii) 180°

(iv) 109°

14. The electronic configurations of three elements, A, B and C are given below. Answer the questions 14 to 17 on the basis of these configurations.

A $1s^2 2s^2 2p^6$

B $1s^2 2s^2 2p^6 3s^2 3p^3$

C $1s^2 2s^2 2p^6 3s^2 3p^5$

15. Stable form of A may be represented by the formula :

(xxxiii) A

(xxxiv) A_2

(xxxv) A_3

(xxxvi) A_4

16. Stable form of **C** may be represented by the formula :

(xxxvii) C

17. The molecular formula of the compound formed from B and C will be

(i) BC

(ii) B_2C

(iii) BC_2

(iv) BC_3

18. The bond between B and C will be

(v) Ionic

(vi) Covalent

(vii) Hydrogen

(viii) Coordinate

19. Which of the following order of energies of molecular orbitals of N_2 is correct?

(i) $(\pi 2p_y) < (\sigma 2p_z) < (\pi^* 2p_x) \approx (\pi^* 2p_y)$

(ii) $(\pi 2p_y) > (\sigma 2p_z) > (\pi^* 2p_x) \approx (\pi^* 2p_y)$

(iii) $(\pi 2p_y) < (\sigma 2p_z) > (\pi^* 2p_x) \approx (\pi^* 2p_y)$

(iv) $(\pi 2p_y) > (\sigma 2p_z) < (\pi^* 2p_x) \approx (\pi^* 2p_y)$

20. Which of the following statement is **not** correct from the view point of molecular orbital theory?

(ix) Be_2 is not a stable molecule.

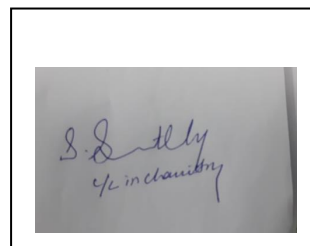
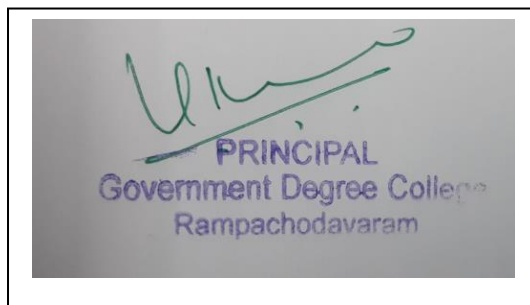
(x) He_2 is not stable but He_2^+ is expected to exist.

(xi) Bond strength of N_2 is maximum amongst the homonuclear diatomic molecules belonging to the second period.

(xii) The order of energies of molecular orbitals in N_2 molecule is

$$\sigma 2s < \sigma^* 2s < \sigma 2p_z < (\pi 2p_x = \pi 2p_y) < (\pi^* 2p_x = \pi^* 2p_y) < \sigma^* 2p_z$$

S.NO	Name of the student	CLASS	SIGNATURE OF THE STUDENT
1	K.APPANNA BABU	IIIMPC	K.APPANNA BABU
2	L.RAMA DEVI	IIIMPC	L.RAMA DEVI
3	P.NARSAYYA	IIIMPC	P.NARSAYYA
4	V.BHEEMI REDDY	IIIMPC	V.BHEEMI REDDY
5	B.V.KRISHNA REDDY	IIIMPC	B.V.KRISHNA REDDY
6	M.VENKATA LAKSHMI	IIIMPC	M.VENKATA LAKSHMI
7	S.KARNAMMA	IIIMPC	S.KARNAMMA
8	G.SAI PRASAD	IIIMPC	G.SAI PRASAD
9	K.VEERA DURGA RAO	IIIMPC	K.VEERA DURGA RAO
10	M.RAMI REDDY	IIIMPC	M.RAMI REDDY
11	Y.SATYANARAYANA REDDY	IIMPC	Y.SATYANARAYANA REDDY
12	B.GANGA BHAVANI	IIMPC	B.GANGA BHAVANI
13	CH.BYULAKSHMI	IIMPC	CH.BYULAKSHMI
14	CH.SATYAVATHI	IIMPC	CH.SATYAVATHI
15	T.LOVA KUMARI	IIMPC	T.LOVA KUMARI
16	S.PAVAN KALYAN	IIMPC	S.PAVAN KALYAN
17	S.SAI DURGA	IIMPC	S.SAI DURGA
18	S.SIVA NANDHINI	IIMPC	S.SIVA NANDHINI
19	P.ARUNA	II MPC	P.ARUNA
20	J.SUBHADRA	IMPC	J.SUBHADRA
21	K.ANUSHA	IMPC	K.ANUSHA
22	K.PRAKASH REDDY	IMPC	K.PRAKASH REDDY
23	K.RAJA BABU	IMPC	K.RAJA BABU
24	M.ANIL KUMAR DORA	IMPC	M.ANIL KUMAR DORA



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REPORT ON STUDENT QUIZ –I,II,III- B.Sc

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Majority of the students from I B.Sc II B.Sc III B.Sc were attended for the QUIZ .The topic was enlightened and informative to the students.

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RAMPACHODAVARAM,
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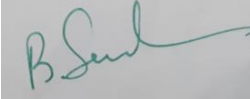
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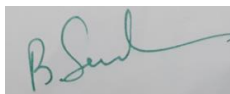
PRINCIPAL
Government Degree College
Rampachodavaram



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DEPARTMENT OF CHEMISTRY

CIRCULAR

The Department of Chemistry is going to be conducted a student Quiz on 15-02-2019. All the students must attend the Quiz in The chapter Group – 13 Elements Chemistry Lab Room:17 without fail

A rectangular box containing a handwritten signature in blue ink. The signature appears to be 'B. Surendra'.

B.SURENDRA

Lecturer in Chemistry

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LECTURE IN CHEMISTRY



Quiz

This set of Chemistry Multiple Choice Questions & Answers (MCQs) focuses on “p-Block Elements – Group 13 Elements: Boron Family”.

1. Are group 13 elements a part of p block elements?

- a) Yes
- b) No
- c) Only a few
- d) Only one

Answer: a

2. Which of the following group's elements have smaller atomic radii?

- a) Group 1 elements
- b) Group 2 elements
- c) Group 13 elements
- d) All have the same atomic radii

Answer: c

3. The atomic radius of gallium is greater than that of aluminum.

- a) True
- b) False

Answer: b

4. Gallium remains liquid up to _____ Kelvin.

- a) 2176
- b) 2376
- c) 2476
- d) 2276

Answer: d

5. The ionization enthalpy _____ down the group in the family.

- a) Increases
- b) Decreases
- c) Constant
- d) Is a regular

View Answer

Answer: d

6. Inert pair affect _____ down the group.

- a) Increases
- b) Decreases
- c) Constant
- d) Is a regul

Answer: a

7. Which of the following element exhibits + 3 Oxidation State only?

- a) Gallium
- b) Thallium
- c) Indium
- d) Aluminium

Answer:d

8. Which of the following is true regarding reducing character?

- a) Gallium < aluminium > indium > thallium
- b) Aluminium > gallium > indium > thallium
- c) Aluminium > gallium < indium > thallium
- d) Gallium > aluminium > indium > thallium

Answer: b

9. Complex formation is more likely to be possible in _____

- a) alkali metals
- b) alkaline earth metals
- c) boron family
- d) equally likely

Answer: c

10. The compounds formed by the Boron family are _____

- a) ionic
- b) covalent
- c) both ionic and covalent
- d) neither ionic nor covalent

Answer:c

11. What is the chemical formula of aluminium carbide?

- a) AlC
- b) AlC₃
- c) AlC₂
- d) AC₃

View Answer

Answer: b

Explanation: 4 moles of aluminium atom combines with 3 moles of carbon atom on heating, in order to form aluminium carbide. Aluminium carbide is ionic in nature and it also forms methane with water. Its chemical formula is given by AlC₃.

12. When boron reacts with nitrogen which of the following compound is formed?

- a) Boron oxide
- b) Boron nitrate
- c) Boron hydrides
- d) Boron nitride

View Answer

Answer: d

Explanation: On heating, two moles of boron atom combine with one mole of a nitrogen molecule in order to form 2 moles of boron nitride. Aluminium also when reacted with nitrogen forms aluminium nitride in the same way.

13. What forms when boron combines with caustic soda?

- a) Formation of oxygen
- b) Formation of washing soda
- c) Formation of Boron nitride
- d) Formation of sodium borate

View Answer

Answer: d

Explanation: Two moles of boron atoms fuse with 6 moles of sodium hydroxide in order to form 2 moles of sodium borate and three moles of hydrogen molecules. Sodium hydroxide is also known as caustic soda.

14. The metallic character of _____ is less than that of alkaline earth metals.

- a) Boron family
- b) Alkali metals
- c) Magnesium
- d) Hydrogen

[View Answer](#)

Answer: a

Explanation: The elements of the Boron family are less electropositive than the alkaline earth metals due to their smaller size and higher ionization enthalpies. On moving down the group, the electropositive character first increases from Boron to aluminium and then decreases from gallium so thallium due to the presence of d and f orbitals which causes poor shielding.

15. Which of the following is the correct order for the stability of plus one oxidation State?

- a) $Ga < In < Tl$
- b) $Ga < In > Tl$
- c) $Ga > In < Tl$
- d) $Ga > In > Tl$

[View Answer](#)

Answer: a

Explanation: The correct order of the increased stability of + 1 oxidation state in gallium, indium and thallium is that; gallium stability is less than that of indium and indium stability is less than that of thallium.

16. What is the atomic number of thallium?

- (a) 81
- (b) 77
- (c) 66
- (d) None of the above

Answer: (a) The atomic number of thallium is 81.

17. Which of the following element primarily shows a +1 oxidation state?

- (a) Boron
- (b) Aluminium
- (c) Thallium
- (d) None of the above

Answer: (c) Thallium primarily shows a +1 oxidation state.

18. Which of the following electrolyte is used to extract aluminium?

- (a) Pure alumina with bauxite and molten cryolite
- (b) Fused cryolite with fluorspar
- (c) Pure alumina in molten cryolite
- (d) Fused cryolite with felspar

Answer: (c) Pure alumina in molten cryolite is used to extract aluminium.

19. The electrolytic refining of aluminium is known as

- (a) Hoop's process
- (b) Baeyer's process
- (c) Serpeck's process
- (d) None of the above

Answer: (a) The electrolytic refining of aluminium is known as Hoop's process.

20. Which of the following is the hardest compound of boron?

- (a) Boron carbide
- (b) Boron fluoride
- (c) Boron nitride
- (d) None of the above

Answer: (c) Boron nitride is the hardest compound of boron.

21. How does the oxidation state varies when we move from boron to thallium?

Answer: When we move from boron to thallium, the stability of the +3 oxidation state decreases while that of the +1 oxidation state increases.

22. Why is BCl_3 more than stable $TiCl_3$?

Answer: BCl_3 is more than stable $TiCl_3$ because, in boron ($2s^2, 2p^1$), all three valence electrons are available for bonding with chlorine. While in $TiCl_3$ ($6s^2, 6p^1$), the 6s orbital experiences a maximum screening effect. Thus, only 6 p^1 electron is available for bonding with chlorine, making BCl_3 more than stable $TiCl_3$.

23. What happens when we heat boric acid?

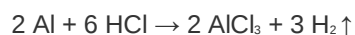
Answer: Boric acid (H_3BO_3) forms metaboric acid (HBO_2) when heated above 370K, while on further heating, it oxidises to boric oxide (B_2O_3)



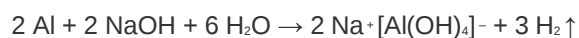
24. How can you prove the amphoteric nature of aluminium?

Answer: We can justify the amphoteric nature of aluminium by reacting it with an acid and a base.

Reaction with an Acid: Aluminium reacts with an acid to yield salt and hydrogen gas.



Reaction with a base: Aluminium reacts with a base to yield salt and hydrogen gas.



25. Why is the bond length of the B-F bond longer in BF_4^- while shorter in BF_3 ?

Answer: The bond length of the B-F bond is longer in BF_4^- while shorter in BF_3 because the BF_4^- is sp^3 hybridised while BF_3 is sp^2 hybridised. An sp^3 hybridised orbital is longer than the sp^2 hybridised orbital. Thus, the bond length of the B-F bond is longer in BF_4^- while shorter in BF_3 .

26. Why do boron halides form addition compounds with ammonia and amines?

Answer: Boron halides are electron deficient in nature, so they form additional compounds with electron-rich species like ammonia and amines.

27. Why does gallium have large ionisation enthalpy than aluminium?

Answer: Gallium has larger ionisation enthalpy than aluminium because gallium has a poor screening effect, due to which the effective nuclear charge in gallium increases.

Thus, gallium has higher ionisation enthalpy than aluminium.

FEED BACK FROM STUDENTS

1. Is the QUIZ useful for you

a) Yes b) No

2. How was the student present Participation

a) Good b) Better C) Excellent

3. Did the event meet your expectation

a) yes b) No

4. Is the content clear by the presenter

a) Yes b) No

5. Did you enjoyed the quiz or felt bored

a) Enjoyed b) Bored

S.NO	NAME OF THE STUDENT	CLASS	SIGNATURE OF THE STUDENT
1	VANJAM DURGA RAO	IBZC	VANJAM DURGA RAO
2	POLISETTY.RAJULAMMA	IBZC	POLISETTY.RAJULAMMA
3	PALLALA.NARAYANA REDDY	IBZC	PALLALA.NARAYANA REDDY
4	MURLA RATNA KUMARI	IBZC	MURLA RATNA KUMARI
5	MADAKAM.VAMSI KUMAR	IMPC	MADAKAM.VAMSI KUMAR
6	MIRTHIVADA.RAJA REDDY	IMPC	MIRTHIVADA.RAJA REDDY
7	PALLALA.RAJYA LAKSMI	IMPC	PALLALA.RAJYA LAKSMI
8	BOGGULA. RAVI KIRAN	IMPC	BOGGULA. RAVI KIRAN
9	B.ABBAI REDDY	IIMPC	B.ABBAI REDDY
10	G.SRINIVASA RAO	IIMPC	G.SRINIVASA RAO
11	J.SUBHADRA	IIMPC	J.SUBHADRA
12	K.CHINNA REDDY	IIMPC	K.CHINNA REDDY
13	K.PRAKASH REDDY	IIMPC	K.PRAKASH REDDY
14	V.PRABHAVATHI	IIMPC	V.PRABHAVATHI
15	K.SATYANARAYANA REDDY	IIMPC	K.SATYANARAYANA REDDY
s16	M.ANIL KUMAR DORA	IIMPC	M.ANIL KUMAR DORA
17	Y.SATYANARAYANA REDDY	IIIMPC	Y.SATYANARAYANA REDDY
18	B.GANGA BHAVANI	IIIMPC	B.GANGA BHAVANI
19	CH.BYULAKSHMI	IIIMPC	CH.BYULAKSHMI
20	CH.SATYAVATHI	IIIMPC	CH.SATYAVATHI
21	T.LOVA KUMARI	IIIMPC	T.LOVA KUMARI
22	S.PAVAN KALYAN	IIIMPC	S.PAVAN KALYAN
23	S.SAI DURGA	IIIMPC	S.SAI DURGA
24	S.SIVA NANDHINI	IIIMPC	S.SIVA NANDHINI
25	P.ARUNA	IIIMPC	P.ARUNA

GOVERNMENT DEGREE COLLEGE
RAMPACHODAVARAM
ALLURI SEETHA RAMA RAJU DISTRICT

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QUIZ COMPETITION FOR I ,II,III B.Sc

2019-2020

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DEPARTMENT OF CHEMISTRY

REPORT ON STUDENT QUIZ –I,II,III- B.Sc

The Department of Chemistry Conducted a student Seminar for the students of II B.Sc mpc Students on 26-02-2020 .The topic For the QUIZ was Group-13 Elements .

Majority of the students from I B.Sc II B.Sc III B.Sc were attended for the QUIZ .The topic was enlightened and informative to the students.

Few Students have asked questions related to the topic .After the completion of seminar feed back was also Collected from the students.

TO
THE PRINCIPAL,
GOVT.DEGREE COLLEGE,
RAMPACHODAVARAM.

RAMPACHODAVARAM,
24-02-2020.

Respected Sir,

SUB: Seeking permission to conduct student seminars for the year 2019-2020 Req reg.

I B.SURENDRA ,Lecturer in chemistry, Govt.Degree College, Rampachodavaram, planned to conduct student seminars for the year 2019-2020 for I,II,III B.Sc[MPC] classes.

So kindly grant me permission to conduct seminars.

Thanking you sir.

Yours faithfully



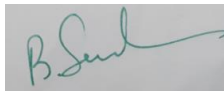
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Government Degree College
Rampachodavaram



GOVERNMENT DEGREE COLLEGE
RAMPACHODAVARAM
DEPARTMENT OF CHEMISTRY

CIRCULAR

The Department of Chemistry is going to be conducted a student Quiz on 26-02-2020. All the students must attend the Quiz in The chapter Group – 13 Elements Chemistry Lab Room:17 without fail

A rectangular box containing a handwritten signature in green ink. The signature appears to be 'B. Surendra'.

B.SURENDRA

Lecturer in Chemistry

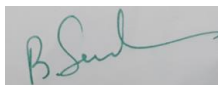
GOVERNMENT DEGREE COLLEGE

RAMPACHODAVRAM

DEPARTMENT OF CHEMISTRY

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B.SURENDRA

LECTURE IN CHEMISTRY

CARBOHYDRATES

Carbohydrates are the most abundant biomolecules on earth. Oxidation of carbohydrates is the central energy-yielding pathway in most non-photosynthetic cells.

Definition: Carbohydrates are polyhydroxy aldehydes or ketones, or substances that yield such compounds on hydrolysis. carbohydrates have the empirical formula $(CH_2O)_n$.

There are three major classes of carbohydrates:

1. Monosaccharides Monosaccharides, or simple sugars, consist of a single polyhydroxy aldehyde or ketone unit. The most abundant monosaccharide in nature is the six-carbon sugar Dglucose, sometimes referred to as dextrose.

2. Oligosaccharides Oligosaccharides consist of short chains of monosaccharide units, or residues, joined by characteristic linkages called glycosidic bonds. The most abundant are the disaccharides, with two monosaccharide units.

Example: sucrose (cane sugar).

3. Polysaccharides The polysaccharides are sugar polymers containing more than 20 or so monosaccharide units, and some have hundreds or thousands of units. Example: starch. Polysaccharides are of two types based on their function and composition. Based on function, polysaccharides are of two types storage and structural.

A. Storage

polysaccharide - starch.

B. Structural polysaccharide - cellulose.

1. Fructose is metabolized by

- A. fructose 1-phosphate pathway
- B. fructose 6-phosphate pathway
- C. glyceraldehyde 3-phosphate pathway
- D. both (a) and (b)

Answer: Option D

2. A common way that cells capture the energy released during the breakdown of large molecules is to add electrons to smaller, specialized molecules that can accept them. This process of electron acceptance is otherwise known as

- A. biosynthesis
- B. metabolism
- C. reduction
- D. catalysis

Answer: Option C

3. Humans are unable to digest

- A. starch
- B. complex carbohydrates
- C. denatured proteins
- D. cellulose

Answer: Option D

4. How many ATP equivalents per mole of glucose input are required for gluconeogenesis?

A. 2

B. 6

C. 8

D. 4

Answer: Option B

5. Which of the following compounds is responsible for coordinated regulation of glucose and glycogen metabolism?

A. NAD^+

B. Fructose 2,6 bisphosphate

C. Acetyl-CoA

D. Fructose 1,6 bisphosphate

Answer: Option B



FEED BACK FROM STUDENTS

1.Is the QUIZ useful for you

a) Yes b) No

2.How was the student present Participation

a)Good b) Better C)Excellent

3.Did the event meet your expectation

a) yes b) No

4.Is the content clear by the presenter

a) Yes b) No

5.Did you enjoyed the quiz or felt bored

a) Enjoyed b) Bored

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8	BOGGULA. RAVI KIRAN	IIMPC	BOGGULA. RAVI KIRAN
9	B.ABBAI REDDY	IIIMPC	B.ABBAI REDDY
10	G.SRINIVASA RAO	IIIMPC	G.SRINIVASA RAO
11	J.SUBHADRA	IIIMPC	J.SUBHADRA
12	K.CHINNA REDDY	IIIMPC	K.CHINNA REDDY
13	K.PRAKASH REDDY	IIIMPC	K.PRAKASH REDDY
14	V.PRABHAVATHI	IIIMPC	V.PRABHAVATHI
15	K.SATYANARAYANA REDDY	IIIMPC	K.SATYANARAYANA REDDY
s16	M.ANIL KUMAR DORA	IIIMPC	M.ANIL KUMAR DORA
17	M.SATYANARAYANA REDDY	IMPC	M.SATYANARAYANA REDDY
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21	V.CHANDRIKA	IMPC	V.CHANDRIKA
22	M.SARASWATHI	IMPC	M.SARASWATHI
23	P.LALITHA	IMPC	P.LALITHA
24	Y.ESWARI	IMPC	Y.ESWARI
25	K.HEMAHAHESWARI	IBZC	K.HEMAHAHESWARI

GOVERNMENT DEGREE COLLEGE
RAMPACHODAVARAM
ALLURI SEETHA RAMA RAJU DISTRICT

DEPARTMENT OF CHEMISTRY

QUIZ COMPETITION FOR I ,II,III B.Sc

2020-2021

**GOVERNMENT DEGREE COLLEGE
RAMPACHODAVARAM**

DEPARTMENT OF CHEMISTRY

REPORT ON STUDENT QUIZ –I,II,III- B.Sc

The Department of Chemistry Conducted a student Seminar for the students of II B.Sc mpc Students on 25-02-2021 .The topic For the QUIZ was Group-13 Elements .

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TO
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RAMPACHODAVARAM.

RAMPACHODAVARAM,
20-02-2021.

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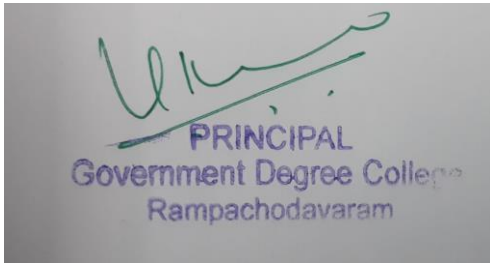
SUB: Seeking permission to conduct student seminars for the year 2020-2021 Req reg.

I B.SURENDRA ,Lecturer in chemistry, Govt.Degree College, Rampachodavaram, planned to conduct student seminars for the year 2020-2021 for I,II,III B.Sc[MPC] classes.

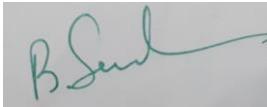
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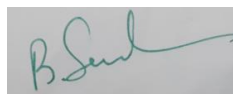
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Lecturer in Chemistry

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RAMPACHODAVRAM

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B.SURENDRA

LECTURE IN CHEMISTRY



Chemistry Questions and Answers – p-Block Elements – Group 14 Elements: Carbon Family

This set of Chemistry Multiple Choice Questions & Answers (MCQs) focuses on “p-Block Elements – Group 14 Elements: Carbon Family”.

1. The ionization enthalpy and density increase in the group from top to bottom.
- a) true
 - b) false

Answer: b

2. Why hydrides of Germanium are known as _____
- a) silanes
 - b) germanes
 - c) stannum
 - d) plumbane
- Answer: b

3. The group 14 elements form _____ hydrides.
- a) metallic
 - b) ionic

- c) covalent
- d) both covalent and

Answer: c

5. Which of the following group 14 elements is a metal?

- a) Stannum
- b) Carbon
- c) Germanium
- d) Silicon

Answer: a

6. What is the colour of silicon?

- a) blue
- b) silver
- c) black
- d) light brown

Answer: d

7. What is the Fajan's rule about?

- a) electronegativity
- b) ionic compounds
- c) Oxidation State
- d) covalent compounds

Answer: c

8. Do Carbon family elements show multiple bonding?

- a) Yes
- b) Maybe
- c) No
- d) Cannot say

Answer: a

9. All the elements in group 14 exhibit tetravalency.

- a) true
- b) false

Answer: a

10. Is catenation possible in carbon?

- a) Yes
- b) Maybe
- c) No
- d) Cannot say

Answer: a

FEED BACK FROM STUDENTS

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- a) Yes
- b) No

2.How was the student present Participation

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- C)Excellent

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GOVERNMENT DEGREE COLLEGE
RAMPACHODAVARAM
ALLURI SEETHA RAMA RAJU DISTRICT

DEPARTMENT OF CHEMISTRY

QUIZ COMPETITION FOR I ,II,III B.Sc

2021-2022

**GOVERNMENT DEGREE COLLEGE
RAMPACHODAVARAM
DEPARTMENT OF CHEMISTRY**

TO

RAMPACHODAVARAM

THE PRINCIPAL

11-02-2022

**GOVERNMENT DEGREE COLLEGE
RAMPACHODAVARAM**

Sir,

Sub: Application submitted for seeking permission to conduct quiz programme req reg.

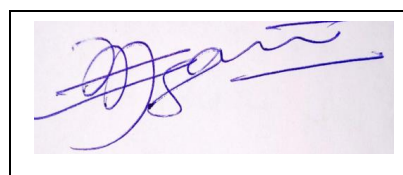
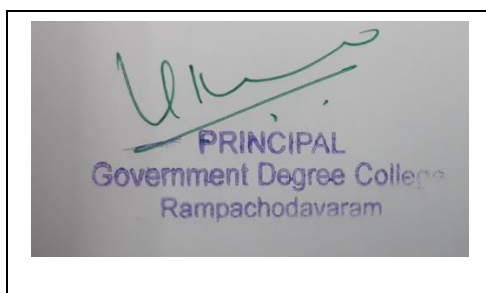
On the occasion of celebration of National Science day, we are planning to conduct a quiz programme to i,ii,&iii BSC MPC&BZC students.

This is for your kind information. So kindly grant me permission to conduct quiz programme on 17/02/2022.

Thanking you sir

Yours faithfully

M. GANESH



GOVERNMENT DEGREE COLLEGE

RAMPACHODAVRAM

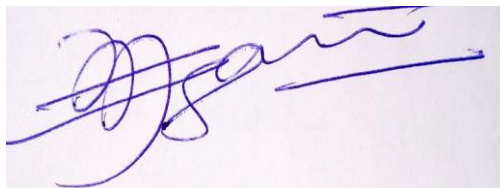
DEPARTMENT OF CHEMISTRY

CIRCULAR

The Department of Chemistry is planned to conduct an inter group quiz programme for all the three years of BSc MPC&BZC Students on 17-02-2022. I Cordially invite all the staff members to attend quiz programme. All the students must attend the quiz in Room No-10 without fail.

M GANESH

LECTURER IN CHEMISTRY



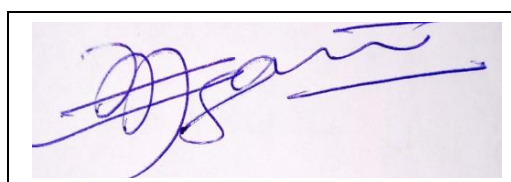
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LECTURE IN CHEMISTRY

**GOVERNMENT DEGREE COLLEGE,
RAMPACHODAVARAM**

DEPARTMENT OF CHEMISTRY

QUIZ- BSC[MPC &BZC]

Chemistry MCQ Questions and Answers

1. The nucleus of an atom consists of _____

- A. electrons and protons
- B. electrons and neutrons
- C. protons and neutrons
- D. All of the above

Answer: protons and neutrons

2. The most electronegative element among the following is _____

- A. bromine
- B. sodium
- C. fluorine
- D. oxygen

Answer: fluorine

3. The metal used to recover copper from a solution of copper sulphate is _____

- A. Ag
- B. Na
- C. Hg
- D. Fe

Answer: Fe

4. The main use of salt in the diet is to _____

- A. produce in small amounts the hydrochloric acid required for the digestion of food
- B. make the taste of food better
- C. increase the solubility of food particles in water
- D. ease the process of cooking

Answer: produce in small amounts the hydrochloric acid required for the digestion of food

5. The mass of one Avogadro number of helium atom is _____

- A. 4.00 gram

B. $4 \times 6.02 \times 10^{23}$ gram

C. 8.00 gram

D. 1.00 gram

Answer: $4 \times 6.02 \times 10^{23}$ gram

6. What are the number of moles of Co_2 which contains 16 g of oxygen?

A. 0.25 mole

B. 0.4 mole

C. 0.2 mole

D. 0.5 mole

Answer: 0.5 moles

7. The number of water molecules present in a drop of water (volume 0.0018 ml) at room temperature is

A. 6.023×10^{23}

B. 4.84×10^{17}

C. 6.023×10^{19}

D. 1.568×10^3

Answer: 6.023×10^{19}

8. The most abundant rare gas in the atmosphere is _____

A. Ne

B. He

C. Ar

D. Xe

Answer: Ar

9. The Latin word Formica means ant. The name formic acid is derived from this Latin word because _____

A. ants are attracted by the odour of this acid

- B. this acid was first obtained by the distillation of ants
- C. this acid, in ancient times, was used to eliminate ant-hills
- D. this corrosive acid is secreted by ants to drive away their enemies

Answer: this acid was first obtained by the distillation of ants

10. The inherited traits of an organism are controlled by _____

- A. RNA molecules
- B. enzymes
- C. DNA molecules
- D. nucleotides

Answer: DNA molecules

E. bauxite

11. magnetite The ore which is found in abundance in India is____

- A.
- B. fluorspar
- C. monazite

Answer: monazite

12. The heat energy produced when the human body metabolises 1 gram of fat is__

- A. 30 KJ
- B. 1 KJ
- C. 39 KJ
- D. 29 KJ

Answer: 39 KJ

13. The mass number of a nucleus is_____

- A. the sum of the number of protons and neutrons present in the nucleus

- B. always more than the atomic weight
- C. always less than its atomic number
- D. a fraction

Answer: the sum of the number of protons and neutrons present in the nucleus

14. The material which can be deformed permanently by heat and pressure is called _____

- A. thermoplastic
- B. polymer
- C. thermoset
- D. chemical compound

Answer: thermoset

15. The items amenable to detection by soft x-rays are _____

- A. genuine coins from counterfeit coins
- B. narcotics
- C. lead in bullets
- D. contrabands

Answer: genuine coins from counterfeit coins

16. The Most commonly used bleaching agent is _____

- A. alcohol
- B. chlorine
- C. carbon dioxide
- D. sodium chlorine

Answer: chlorine

17. The nucleus of a hydrogen atom consists of _____

- A. 1 neutron only
- B. 1 proton only

C. 1 proton + 2 neutron

D. 1 electron only

Answer: 1 proton only

18. The metal that is used as a catalyst in the hydrogenation of oils is_____

A. Pb

B. Ni

C. Pt

D. Cu

Answer: Ni

19. The main buffer system of the human blood is_____

A. H_2CO_3 - HCO_3

B. H_2CO_3 - CO_3^{2-}

C. CH_3COOH - CH_3COO

D. NH_2CONH_2 - NH_2CONH^+

Answer: H_2CO_3 - HCO_3

20. The heat required to raise the temperature of a body by 1 K is called _____

A. thermal capacity

B. specific heat

C. water equivalent

D. None of the above

Answer: thermal capacity

21. The nuclear particles which are assumed to hold the nucleons together are_____

A. positrons

B. neutrons

C. electrons

D. mesons

Answer: mesons

22. The octane number of zero is assigned to _____

A. n-heptane

B. 2-methyl octane

C. iso-octane

D. 3-methyl octane

Answer: n-heptane

FEED BACK FROM STUDENTS

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2	A.SANTHOSH	I MPC	A.SANTHOSH
3	A.VENKATESULU	I MPC	A.VENKATESULU
4	K.SAI TEJA SRI	I MPC	K.SAI TEJA SRI
5	D.PAVAN	I MPC	D.PAVAN
6	K.HARI CHARAN REDDY	I MPC	K.HARI CHARAN REDDY
7	MAMIDI.CHINNAMMI	I MPC	MAMIDI.CHINNAMMI
8	K.K.BHAVANI	II MPC	K.K.BHAVANI
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